

Bis(2-formylphenolato- $\kappa^2 O,O'$)manganese(II)

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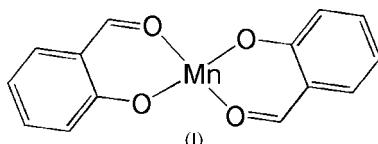
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The title compound, $[\text{Mn}(\text{C}_7\text{H}_5\text{O}_2)_2]$, is a mononuclear manganese(II) complex. The Mn^{II} atom, lying on an inversion centre, is four-coordinated by four O atoms from two salicylaldehyde ligands, forming a square-planar geometry.

Comment

Manganese(II) complexes are very important in biological chemistry and supramolecular chemistry (Miyasaka *et al.*, 1996; Ciringh *et al.*, 1997; Mabad *et al.*, 1986). We report here the crystal structure of the new title manganese(II) complex, (I).



The Mn^{II} ion in complex (I), lying on an inversion centre, is four-coordinated by four O atoms from two salicylaldehyde ligands, forming a square-planar geometry (Fig. 1). The bond lengths and angles (Table 1) involving the Mn^{II} ion are comparable with the values observed in other manganese complexes (Nakasuka *et al.*, 1985; Zhang, 2006; Gao & Liu, 2005; Okabe & Koizumi, 1998).

Experimental

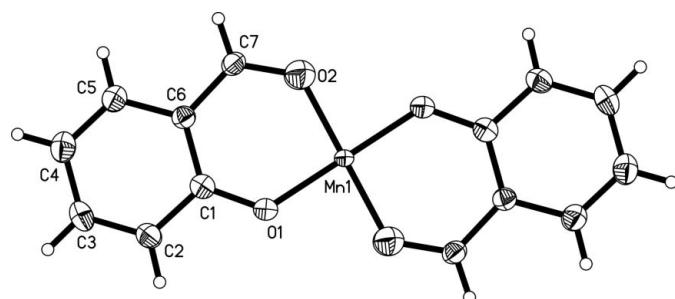
Salicylaldehyde (1.0 mmol, 122.1 mg) and $\text{Mn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$ (0.5 mmol, 104.5 mg) were dissolved in ethanol (80 ml). The mixture was refluxed at 345 K under an argon atmosphere for about 1 h to give a red solution. After allowing this solution to stand in air for 7 d, deep-brown plate-shaped crystals were formed at the bottom of the vessel.

Crystal data

$[\text{Mn}(\text{C}_7\text{H}_5\text{O}_2)_2]$	$Z = 2$
$M_r = 297.16$	$D_x = 1.625 \text{ Mg m}^{-3}$
Monoclinic, $P2_{\bar{1}}/c$	Mo $K\alpha$ radiation
$a = 12.918 (2) \text{ \AA}$	$\mu = 1.09 \text{ mm}^{-1}$
$b = 5.831 (1) \text{ \AA}$	$T = 293 (2) \text{ K}$
$c = 8.101 (3) \text{ \AA}$	Plate, brown
$\beta = 95.54 (3)^\circ$	$0.10 \times 0.10 \times 0.03 \text{ mm}$
$V = 607.4 (3) \text{ \AA}^3$	

Data collection

Bruker SMART CCD area-detector diffractometer	5037 measured reflections
ω scans	1446 independent reflections
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	1133 reflections with $I > 2\sigma(I)$
$T_{\min} = 0.899$, $T_{\max} = 0.968$	$R_{\text{int}} = 0.049$
	$\theta_{\max} = 28.3^\circ$

**Figure 1**

The structure of (I), showing 30% probability displacement ellipsoids and the atom-numbering scheme. Unlabelled atoms are related to the labelled atoms by the symmetry operation $(-x, 2 - y, 1 - z)$.

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.065$
 $wR(F^2) = 0.172$
 $S = 1.07$
1446 reflections
88 parameters
H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0704P)^2 + 1.5184P]$$

where $P = (F_o^2 + 2F_c^2)/3$

$$(\Delta/\sigma)_{\max} < 0.001$$

$$\Delta\rho_{\max} = 0.48 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.75 \text{ e } \text{\AA}^{-3}$$

Table 1
Selected geometric parameters (\AA , $^\circ$).

Mn1—O1	1.833 (3)	Mn1—O2	1.844 (4)
O1 ⁱ —Mn1—O1	180	O1—Mn1—O2	94.90 (17)
O1—Mn1—O2 ⁱ	85.10 (17)	O2 ⁱ —Mn1—O2	180
Symmetry code: (i) $-x, -y + 2, -z + 1$.			

All H atoms were placed in idealized positions and constrained to ride on their parent atoms, with C—H distances of 0.93 \AA and with $U_{\text{iso}}(\text{H})$ set to $1.2U_{\text{eq}}(\text{C})$.

Data collection: SMART (Bruker, 1998); cell refinement: SAINT (Bruker, 1998); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 1997a); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997a); molecular graphics: SHELXTL (Sheldrick, 1997b); software used to prepare material for publication: SHELXTL.

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References

- Bruker (1998). SMART (Version 5.628) and SAINT (Version 6.02). Bruker AXS Inc., Madison, Wisconsin, USA.
- Cirighi, Y., Gordon-Wylie, S. W., Norman, R. E., Clark, G. R., Weintraub, S. T. & Horwitz, C. P. (1997). *Inorg. Chem.* **36**, 4968–4982.
- Gao, J. & Liu, Y. (2005). *Acta Cryst. E61*, m1692–m1693.
- Mabad, B., Cassoux, P., Tuchagues, J.-P. & Hendrickson, D. N. (1986). *Inorg. Chem.* **25**, 1420–1431.
- Miyasaka, H., Matsumoto, N., Okawa, H., Re, N., Gallo, E. & Floriani, C. (1996). *J. Am. Chem. Soc.* **118**, 981–994.
- Nakasuka, N., Azuma, S., Katayama, C., Honda, M., Tanaka, J. & Tanaka, M. (1985). *Acta Cryst. C41*, 1176–1179.
- Okabe, N. & Koizumi, M. (1998). *Acta Cryst. C54*, 288–290.
- Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
- Sheldrick, G. M. (1997a). SHELXS97 and SHELXL97. University of Göttingen, Germany.
- Sheldrick, G. M. (1997b). SHELXTL. Version 5.10. Bruker AXS Inc., Madison, Wisconsin, USA.
- Zhang, Y.-L. (2006). *Acta Cryst. E62*, m900–m901.